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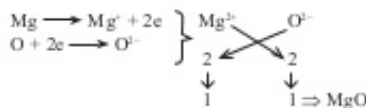
Explanation / Penerangan:

- (a) Electron arrangement of sodium atom is 2.8.1. Sodium atom has one valence electron. Therefore sodium atom is not stable. Sodium atom releases one electron to achieve a stable octet electron arrangement to form sodium ion, Na^+ with electron arrangement 2.8.
Susunan elektron atom natrium ialah 2.8.1. Atom natrium mempunyai satu elektron valens. Dengan itu atom natrium tidak stabil. Atom natrium melepaskan satu elektron ini untuk mencapai susunan elektron oktet yang stabil membentuk ion natrium, Na^+ dengan susunan elektron 2.8.
- (b) Electron arrangement of chlorine atom is 2.8.7. Chlorine atom has seven valence electrons. Chlorine atom receives one electron to achieve stable octet electron arrangement to form chloride ion, Cl^- with an octet arrangement of electron 2.8.8.
Susunan elektron bagi atom klorin ialah 2.8.7. Atom klorin mempunyai tujuh elektron valens. Atom klorin mencapai susunan elektron oktet yang stabil dengan menerima satu elektron membentuk ion klorida, Cl^- dengan susunan elektron 2.8.8.
- (c) Sodium ions, Na^+ and chloride ions, Cl^- ions are attracted with strong electrostatic force. The bond formed is called ionic bond. / Ion natrium, Na^+ dan ion klorida, Cl^- ditarik dengan daya elektrostatik yang kuat. Ikatan yang terbentuk dinamakan ikatan ion.

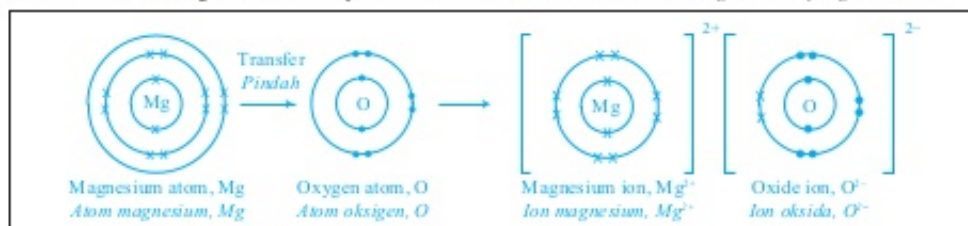
(ii) **Magnesium oxide / Magnesium oksida**

Predict the formula / Ramal formula:

Element Unsur	Proton number Nombor proton	Electron arrangement Susunan elektron
Mg	12	2.8.2
O	8	2.6



Draw the electron arrangement of the compound formed. / Lukiskan susunan elektron bagi sebatian yang terbentuk.



Explanation / Penerangan:

- (a) The electron arrangement of magnesium atom is 2.8.2. Magnesium atom has two valence electrons. Therefore magnesium atom is not stable. Magnesium atom releases two valence electrons to achieve a stable octet electron arrangement to form magnesium ion, Mg^{2+} with electron arrangement 2.8.
Susunan elektron atom magnesium 2.8.2. Atom magnesium mempunyai dua elektron di petala terluar. Maka atom magnesium tidak stabil. Atom magnesium melepaskan dua elektron valens untuk mencapai susunan elektron oktet yang stabil membentuk ion magnesium, Mg^{2+} dengan susunan elektron 2.8.
- (b) The electron arrangement of oxygen atom is 2.6. Oxygen atom is also unstable. Oxygen atom receives two electrons to achieve a stable octet electron arrangement to form oxide ion, O^{2-} with electron arrangement 2.8. / *Susunan elektron atom oksigen ialah 2.6. Atom oksigen juga tidak stabil, atom oksigen menerima dua elektron untuk mencapai susunan elektron oktet yang stabil membentuk ion oksida, O^{2-} dengan susunan elektron 2.8.*
- (c) Magnesium ion, Mg^{2+} and oxide ion, O^{2-} are attracted by strong electrostatic force. The bond formed is called ionic bond. / Ion magnesium, Mg^{2+} dan ion oksida, O^{2-} ditarik dengan daya elektrostatik yang kuat. Ikatan yang terbentuk dinamakan ikatan ion.



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